

**National Exams Dec 2012**  
**04-Chem-A4, Chemical Reactor Engineering**  
**3 hours duration**

**NOTES:**

1. If doubt exists as to the interpretation of any question, the candidate is urged to submit with the answer paper, a clear statement of any assumptions made.

2. This is an OPEN BOOK EXAM.  
Any non-communicating calculator is permitted.

You may bring to this examination

- a) Any edition of the Fogler Reaction Engineer text (annotated in margins, flagged, etc. as desired).
- b) your own unit conversion tables and/or mathematical tables (such as a CRC Handbook).
- c) a non-communicating, programmable electronic calculator including a small operating guide. Please indicate the name and model of your calculator on the first inside left-hand sheet of the exam work book.  
Graph paper will be provided.

3. Four (4) questions constitute a complete exam paper. Complete 4 of 5 questions. Clearly indicate which questions are to be marked.

4. Each question is of equal value.

5. Most questions require an answer in problem solution format. Clarity and organization of the answer are important. Technical content is the key factor in the answers. Thus it is expected that you clearly outline and document your problem solution strategy. All equations used must be clearly documented prior to the use of specific values for the variables, and the parameters must be identified with appropriate units and significant figures. Citation of the origin of significant formulas used (e.g., Fogler, eq. (3-44)) would be of value to the examiner. Please treat all the data provided as exact and give your answers to three significant figures.

6. Each question worth is 20 marks. Marking schemes are indicated by the numbers in brackets such as "{5 marks}" following each part of a problem. Total for the exam is 80 marks. Complete 4 of 5 questions.

**Marking Scheme (ONLY complete 4 of 5 questions for a total of 80 marks)**

1. (a) 2 marks, (b) 10 marks, (c) 4 marks, (d) 4 marks
2. (a) 6 marks, (b) 6 marks, (c) 2 marks, (d) 6 marks
3. (a) 15 marks, (b) 5 marks
4. (a) 10 marks, (b) 10 marks
5. (a) 8 marks, (b) 8 marks, (c) 4 marks

1. A first-order, homogeneous, liquid-phase reaction ( $A \rightarrow B$ ) occurs in a CSTR that operates adiabatically at steady state. The reaction is exothermic. Data are as follows and should be considered as exact quantities:

$$C_A^\circ = 1.75 \text{ mol}_A/\text{L} \text{ (Inlet concentration of A)}$$

$$C_p = 2.5 \text{ J/g}\cdot\text{K} \text{ (The heat capacity of the reacting fluid; assumed constant at all conversions.)}$$

$$\Delta H_r = -27\,000 \text{ J/mol}_A$$

$$q^\circ = 30 \text{ L/min} \text{ (Flow into the CSTR.)}$$

$$(-r_A) = 1.8(10^{13})e^{(-84000/RT)}C_A \text{ (The rate of reaction [mol}_A/\text{L}\cdot\text{s]}; R = 8.314 \text{ [J/mol}\cdot\text{K]}; T \text{ is in [K].)}$$

$$T^\circ = 315 \text{ K} \text{ (Inlet temperature of the feed.)}$$

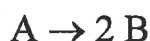
$$V = 28 \text{ L} \text{ (Reactor volume)}$$

$$X_A = 0.9 \text{ (Conversion of A leaving the reactor.)}$$

$$\rho = 500 \text{ g/L} \text{ (Liquid density, assumed constant throughout the reaction.)}$$

- What are the units of the constant  $1.8(10^{13})$  in the rate equation? {2 marks}
- Calculate the temperature at which this reactor operates under steady operating conditions. {10 marks}
- Calculate the rate constant for conditions within the CSTR. {4 marks}
- Calculate the rate of heat evolved within this reactor (in J/s) during steady-state operation. {4 marks}

2. The reaction of pure diphosgene  $\text{ClCOOCCl}_3$  (say A) occurs in the gas phase to produce two molecules of  $\text{COCl}_2$  (say B). At a constant temperature of  $280^\circ\text{C}$ , the stoichiometry for the reaction is:



Total pressure data for this reaction are provided in the table below.

Time, s	Total pressure, torr
0	15
500	18.9
800	20.7
1300	23
1800	24.8

- By plotting the data to the integrated form of a second order reaction, show that this model fails to fit the data. (Graphical or tabular solutions are equally acceptable.) {6 marks}
- Show that a first-order rate law fits the data. {6 marks}
- Report the mean rate constant with appropriate units. {2 marks}
- If reducing the temperature to  $255^\circ\text{C}$  causes the rate constant to decrease to half the value obtained in part (c), what would be the activation energy for this process in  $\text{kJ/mol}$ ? {6 marks}

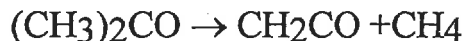
**3. A problem on adiabatic conversion in a plug flow reactor**

- [15] a) Gaseous reactant A is flowing into an adiabatic plug flow reactor at 0.0020 kmol/s carried by steam entering at 0.020 kmol/s. The (absolute) pressure of the combined stream is 105 kPa. The heat of reaction of A is -150 kJ/mol (i.e., is exothermic). The heat capacity of steam is 2.3 kJ/kg•K, a value that may be assumed to describe the heat capacity of the entire stream. Also, the molar mass of the stream may be assumed to be that of steam. The temperature of the feed is exactly 600 K. At that temperature, the rate constant for the reaction is 0.3 s<sup>-1</sup>. The energy of activation for the reaction is 80,000 kJ/mol.

Calculate the temperature of the stream leaving the reactor if 80% of the A has been reacted to product.

- [5] b) How would an increase in pressure of 20% affect the result in part (a)? Please explain.

4. Acetone is pyrolyzed isothermally and irreversibly to ketene in a plug flow reactor at 520°C and 1.00 atm total pressure:



The reactor used in this study has a 3.3 cm inside diameter and 80 cm length. Molar mass of acetone is 58 g/mol. The table below provides conversions ( $X_A$ ) for a feed of pure acetone to the reactor.

Feed Flow, g/h	$X_A$
126	0.05
46	0.13
21	0.24
12	0.35

- [10] a) Show that a first-order kinetic model is reasonable for this reaction and calculate the rate constant in units of 1/h.
- [10] b) Check your result for  $k$  by means of a "differential reactor" analysis, in which you assume

$$\Delta V / \Delta X \approx F_{A^0} / (-r_A)$$

for changes in concentration that are differential (by convention less than about 10%). For such low conversions, it is permissible to neglect volume changes due to stoichiometry.

In this equation,  $\Delta V$  is the volume of the reactor,  $\Delta X$  is the conversion of reactant A,  $F_{A^0}$  is the molar feed rate of reactant A, and  $(-r_A)$  is the rate of disappearance of A by chemical reaction when the concentration of A is close to its value in the feed stream. (The dimensions of both sides of the equation are therefore those of volume.)

5. At a temperature of 25°C, a mixed ethanol-water solvent is 0.0677 molar in formic acid (HCOOH). Some HCl has been added having been added as a catalyst. The resulting esterification reaction is carried out in a batch reactor and monitored by titration to yield the formic acid concentrations as a function of time that are reported in the table below. Initially no ester is present.



During this process, the ethanol concentration is in *very large excess* relative to that of formic acid, while the water concentration is *greatly in excess* of that of the ester. Ethanol and ester concentrations are in some constant (though unknown) ratio. Thus effectively, the overall reaction takes place at an unchanging  $[\text{C}_2\text{H}_5\text{OH}]/[\text{H}_2\text{O}]$  ratio.

- [8] a) Show that the reaction is reversible pseudo-first order. For  $\text{A} \leftrightarrow \text{B}$ , one of the integrated forms of the rate expression for such a reaction is given as

$$\ln[(\text{A}-\text{A}_e)/(\text{A}^\circ-\text{A}_e)] = - (k_f + k_r)t$$

where  $\text{A}^\circ$  is the starting concentration of A, and  $\text{A}_e$  is its equilibrium concentration.  $k_f$  and  $k_r$  are the forward and reverse rate constants respectively.

- [8] b) Calculate individual rate constants  $k_f$  and  $k_r$  for the given  $[\text{C}_2\text{H}_5\text{OH}]/[\text{H}_2\text{O}]$  ratio.
- [4] c) Calculate the pseudo-equilibrium constant  $K = k_f/k_r$  for this particular solvent ratio.

t, min	Formic acid, mol/L
0	0.0677
50	0.0615
100	0.0562
160	0.0509
290	0.0428
$\infty$	0.0292

